

## N-Propionylthiourea

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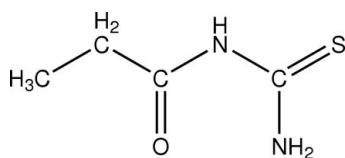
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Key indicators: single-crystal X-ray study;  $T = 298$  K; mean  $\sigma(\text{C}-\text{C}) = 0.004$  Å;  $R$  factor = 0.048;  $wR$  factor = 0.119; data-to-parameter ratio = 17.7.

The molecule of the title compound,  $\text{C}_4\text{H}_8\text{N}_2\text{OS}$ , is essentially planar; it adopts a *trans* configuration with respect to the position of the propionyl group relative to the thiono S atom about the C–N bond. The molecular structure is stabilized by an intramolecular N–H $\cdots$ O hydrogen bond between the propionyl O atom and the amide H atom. Molecules are linked into a two-dimensional network parallel to the (10 $\bar{1}$ ) plane by N–H $\cdots$ O and N–H $\cdots$ S intermolecular hydrogen bonds.

## Related literature

For the crystal structures of thiourea analogues, see: Yusof *et al.* (2007); Rosli *et al.* (2006). For bond-length data, see: Allen *et al.* (1987).



## Experimental

## Crystal data

$\text{C}_4\text{H}_8\text{N}_2\text{OS}$   
 $M_r = 132.19$   
 Monoclinic,  $P2_1/n$   
 $a = 5.0790$  (15) Å  
 $b = 14.342$  (4) Å  
 $c = 9.273$  (3) Å  
 $\beta = 102.744$  (6)°

$V = 658.8$  (3) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.40$  mm<sup>-1</sup>  
 $T = 298$  (2) K  
 $0.48 \times 0.19 \times 0.14$  mm

## Data collection

Bruker SMART APEX CCD area-detector diffractometer  
 Absorption correction: multi-scan (SADABS; Bruker, 2000)  
 $T_{\min} = 0.832$ ,  $T_{\max} = 0.946$

3622 measured reflections  
 1291 independent reflections  
 910 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.029$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.048$   
 $wR(F^2) = 0.119$   
 $S = 1.02$   
 1291 reflections

73 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.23$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.16$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

D–H $\cdots$ A	D–H	H $\cdots$ A	D $\cdots$ A	D–H $\cdots$ A
N2–H2C $\cdots$ O1	0.86	2.00	2.658 (3)	133
N1–H1D $\cdots$ O1 <sup>i</sup>	0.86	2.11	2.935 (3)	160
N2–H2D $\cdots$ S1 <sup>ii</sup>	0.86	2.57	3.409 (3)	166

Symmetry codes: (i)  $x - \frac{1}{2}, -y + \frac{1}{2}, z - \frac{1}{2}$ ; (ii)  $-x, -y, -z + 1$ .

Data collection: SMART (Bruker, 2000); cell refinement: SAINT (Bruker, 2000); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 1997a); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997a); molecular graphics: SHELXTL (Sheldrick, 1997b); software used to prepare material for publication: SHELXTL, PARST (Nardelli, 1995) and PLATON (Spek, 2003).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI2516).

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**supplementary materials**

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## *N*-Propionylthiourea

B. M. Yamin and E. A. Othman

### Comment

Most carbonylthiourea of the type,  $R^1\text{HNC}(\text{S})\text{NHR}^2$ , such as *N*-butanoyl-*N'*-(4-nitrophenyl)thiourea (Yusof *et al.*, 2007) can be prepared from the reaction of carbonylchloride with ammonium thiocyanate which give carbonyl- isothiocyanate, an intermediate for the formation of thiourea moiety when reacted with the amine compounds. However, the title compound (Fig. 1) was unexpectedly obtained when the mixture of propionyl chloride and ammonium thiocyanate was stirred for 30 minutes before adding the amine compound.

The molecule is essentially planar, with a maximum deviation of 0.021 (3) Å for atom C1 from the mean plane. The propionyl group, C1/C2/C3/O1, is *trans* relative to the thiono C4=S1 group across the C4—N1 bond. The bond lengths and angles are in normal ranges (Allen *et al.*, 1987). The molecular structure is stabilized by an intramolecular hydrogen bond, N2—H2C $\cdots$ O1 (Table 1), which forms a S(6) ring. In the crystal structure, the molecules are linked by N1—H1D $\cdots$ O<sup>i</sup> and N2—H2D $\cdots$ S1<sup>ii</sup> intermolecular hydrogen bonds, forming a two-dimensional network (Fig. 2) parallel to the (1 0  $\bar{1}$ ) plane.

### Experimental

A solution of ammonium thiocyanate (0.05 mol, 3.80 g) in acetone (30 ml) was added dropwise to a solution of propionyl chloride (0.05 mol, 4.63 g) in acetone (20 ml). The mixture was stirred for 30 min and the resulting light yellow solution was filtered. Single crystals of the title compound were obtained by slow evaporation of the solution (yield 90%; m.p. 420.2–421.0 K).

### Refinement

H atoms were positioned geometrically (C—H = 0.96 Å and N—H = 0.86 Å) and constrained to ride on their parent atoms, with  $U_{\text{iso}}(\text{H}) = 1.2\text{--}1.5(\text{methyl}) U_{\text{eq}}(\text{parent atom})$ .

### Figures

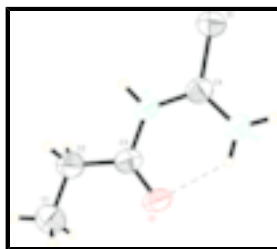


Fig. 1. The molecular structure of the title compound. Displacement ellipsoids are drawn at the 50% probability level. The dashed line indicates a hydrogen bond.

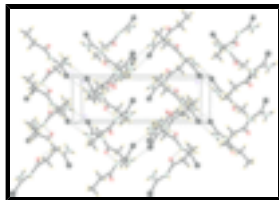


Fig. 2. The molecular packing of the title compound, viewed approximately down the *c* axis. Hydrogen bonds are shown as dashed lines.

## *N*-Propionylthiourea

### Crystal data

$C_4H_8N_2OS$

$M_r = 132.19$

Monoclinic,  $P2_1/n$

Hall symbol:  $-P\ 2_1/n$

$a = 5.0790$  (15) Å

$b = 14.342$  (4) Å

$c = 9.273$  (3) Å

$\beta = 102.744$  (6)°

$V = 658.8$  (3) Å<sup>3</sup>

$Z = 4$

$F_{000} = 280$

$D_x = 1.333$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation

$\lambda = 0.71073$  Å

Cell parameters from 1291 reflections

$\theta = 2.6$ – $26.0$ °

$\mu = 0.40$  mm<sup>-1</sup>

$T = 298$  (2) K

Block, yellow

$0.48 \times 0.19 \times 0.14$  mm

### Data collection

Bruker SMART APEX CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

Detector resolution: 83.66 pixels mm<sup>-1</sup>

$T = 298$ (2) K

$\omega$  scan

Absorption correction: multi-scan (SADABS; Bruker, 2000)

$T_{\min} = 0.832$ ,  $T_{\max} = 0.946$

3622 measured reflections

1291 independent reflections

910 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.029$

$\theta_{\max} = 26.0$ °

$\theta_{\min} = 2.6$ °

$h = -6 \rightarrow 5$

$k = -16 \rightarrow 17$

$l = -9 \rightarrow 11$

### Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.048$

$wR(F^2) = 0.119$

$S = 1.02$

1291 reflections

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0492P)^2 + 0.293P]$$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} = 0.001$

$\Delta\rho_{\max} = 0.23$  e Å<sup>-3</sup>

73 parameters

$$\Delta\rho_{\min} = -0.16 \text{ e } \text{\AA}^{-3}$$

Primary atom site location: structure-invariant direct methods

Extinction correction: none

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.01989 (18)	0.06361 (6)	0.28118 (8)	0.0646 (3)
O1	0.7005 (4)	0.22101 (13)	0.58191 (18)	0.0537 (5)
N1	0.4084 (4)	0.18551 (14)	0.3657 (2)	0.0428 (6)
H1D	0.3630	0.2009	0.2739	0.051*
C4	0.2638 (5)	0.11286 (17)	0.4084 (3)	0.0413 (6)
N2	0.3300 (5)	0.08712 (15)	0.5470 (2)	0.0514 (6)
H2C	0.4587	0.1151	0.6072	0.062*
H2D	0.2448	0.0422	0.5779	0.062*
C3	0.6154 (5)	0.23647 (18)	0.4506 (3)	0.0431 (6)
C2	0.7271 (7)	0.3102 (2)	0.3675 (3)	0.0635 (9)
H2A	0.7959	0.2807	0.2892	0.076*
H2B	0.5808	0.3512	0.3212	0.076*
C1	0.9454 (7)	0.3677 (2)	0.4578 (3)	0.0706 (9)
H1A	1.0051	0.4130	0.3958	0.106*
H1B	1.0940	0.3282	0.5022	0.106*
H1C	0.8785	0.3990	0.5340	0.106*

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
S1	0.0696 (6)	0.0701 (6)	0.0465 (5)	-0.0255 (4)	-0.0038 (4)	0.0050 (4)
O1	0.0653 (13)	0.0566 (12)	0.0324 (10)	-0.0098 (10)	-0.0040 (8)	0.0002 (8)
N1	0.0498 (14)	0.0479 (13)	0.0273 (10)	-0.0060 (11)	0.0013 (9)	0.0028 (9)
C4	0.0445 (15)	0.0408 (15)	0.0385 (14)	0.0040 (12)	0.0089 (11)	-0.0008 (11)
N2	0.0633 (16)	0.0523 (14)	0.0357 (12)	-0.0122 (12)	0.0045 (10)	0.0034 (10)
C3	0.0473 (17)	0.0420 (15)	0.0373 (14)	0.0028 (12)	0.0032 (11)	-0.0015 (11)
C2	0.073 (2)	0.063 (2)	0.0484 (17)	-0.0223 (17)	-0.0002 (15)	0.0073 (14)
C1	0.076 (2)	0.069 (2)	0.064 (2)	-0.0238 (18)	0.0095 (17)	-0.0026 (16)

## supplementary materials

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### Geometric parameters (Å, °)

S1—C4	1.668 (3)	C3—C2	1.492 (4)
O1—C3	1.219 (3)	C2—C1	1.483 (4)
N1—C3	1.377 (3)	C2—H2A	0.97
N1—C4	1.382 (3)	C2—H2B	0.97
N1—H1D	0.86	C1—H1A	0.96
C4—N2	1.308 (3)	C1—H1B	0.96
N2—H2C	0.86	C1—H1C	0.96
N2—H2D	0.86		
C3—N1—C4	128.6 (2)	C1—C2—C3	115.1 (2)
C3—N1—H1D	115.7	C1—C2—H2A	108.5
C4—N1—H1D	115.7	C3—C2—H2A	108.5
N2—C4—N1	117.2 (2)	C1—C2—H2B	108.5
N2—C4—S1	124.4 (2)	C3—C2—H2B	108.5
N1—C4—S1	118.37 (18)	H2A—C2—H2B	107.5
C4—N2—H2C	120.0	C2—C1—H1A	109.5
C4—N2—H2D	120.0	C2—C1—H1B	109.5
H2C—N2—H2D	120.0	H1A—C1—H1B	109.5
O1—C3—N1	122.2 (2)	C2—C1—H1C	109.5
O1—C3—C2	123.6 (2)	H1A—C1—H1C	109.5
N1—C3—C2	114.2 (2)	H1B—C1—H1C	109.5
C3—N1—C4—N2	-0.7 (4)	C4—N1—C3—C2	179.8 (3)
C3—N1—C4—S1	179.7 (2)	O1—C3—C2—C1	-2.5 (5)
C4—N1—C3—O1	0.7 (4)	N1—C3—C2—C1	178.4 (3)

### Hydrogen-bond geometry (Å, °)

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N2—H2C $\cdots$ O1	0.86	2.00	2.658 (3)	133
N1—H1D $\cdots$ O1 <sup>i</sup>	0.86	2.11	2.935 (3)	160
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Symmetry codes: (i)  $x-1/2, -y+1/2, z-1/2$ ; (ii)  $-x, -y, -z+1$ .

Fig. 1

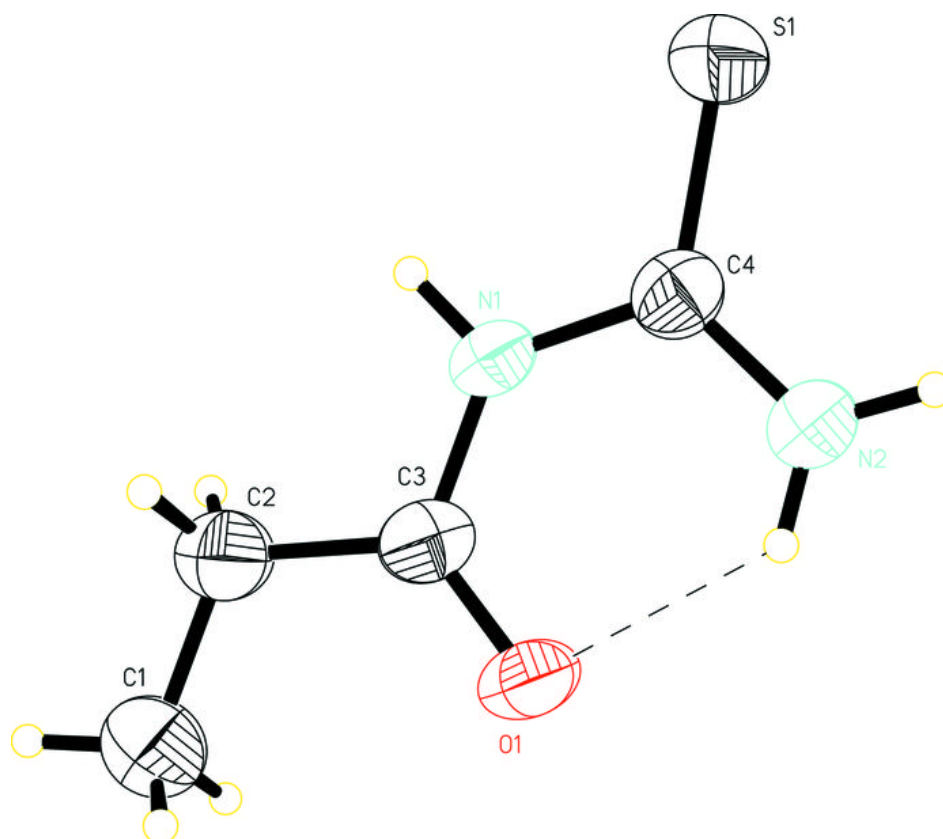


Fig. 2

